



# NCERT Solutions for 12th Class Chemistry: Chapter 15-Polymers



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## NCERT Solutions for 12th Class Chemistry: Chapter 15-Polymers

Class 12: Chemistry Chapter 15 solutions. Complete Class 12 Chemistry Chapter 15 Notes.

### NCERT Solutions for 12th Class Chemistry: Chapter 15-Polymers

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#### EXERCISES (NCERT 12th Chemistry Chapter 15)

**15.1. Explain the terms polymer and monomer.**

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**Ans:** Polymers are high molecular mass substances consisting of a very large number of simple repeating structural units joined together through covalent bonds in a regular fashion. Polymers are also called macromolecules. Some examples are polythene, nylon-66, bakelite, rubber, etc. Monomers are the simple and reactive molecules from which the polymers are prepared either by addition or condensation polymerisation. Some examples are ethene, vinyl chloride, acrylonitrile, phenol and formaldehyde etc.

**15.2. What are natural and synthetic polymers ? Give two examples of each.**

**Ans:**

**1. Natural polymers:** The polymers which occur in nature mostly in plants and animals are called natural polymers. A few common examples are starch, cellulose, proteins, rubber nucleic acids, etc. Among them, starch and cellulose are the polymers of glucose molecules. Proteins are formed from amino acids which may be linked in different ways. These have been discussed in detail in unit 15 on biomolecules. Natural rubber is yet another useful polymer which is obtained from the latex of the rubber tree. The monomer units are of the unsaturated hydrocarbon 2-methyl-1,3-butadiene, also called isoprene.

**Example of natural polymers:** Natural rubber, cellulose, nucleic acids, proteins etc.

**2. Synthetic polymers:** The polymers which are prepared in the laboratory are called synthetic polymers. These are also called man made polymers and have been developed in the present century to meet the ever increasing demand of the modern civilisation.

**Example of synthetic polymers:** Dacron (or terylene), Bakelite, PVC, Nylon-66, Nylon-6 etc.

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**15.3. Distinguish between the terms homopolymer and copolymer and give an example of each.**

**Ans:** Polymers whose repeating structural units are derived from only one type of monomer units are called homopolymers, e.g., PVC polyethene, PAN, teflon, polystyrene, nylon- 6 etc.

Polymers whose repeating structural units are derived from two or more types of monomer molecules are copolymers, e.g., Buna-S, Buna-N, nylon-66, polyester, bakelite.

**15.4. How do you explain the functionality of a monomer?**

**Ans:** Functionality of a monomer implies the number of bonding sites present in it. For example, monomers like propene, styrene, acrylonitrile have functionality of one which means that they have one bonding site.

Monomers such as ethylene glycol, hexamethylenediamine, adipic acid have functionality of two which means that they have two bonding sites.

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**15.5. Define the term polymerisation?**

**Ans:** It is a process of formation of a high molecular Sol. mass polymer from one or more monomers by linking together a large number of repeating structural units through covalent bonds.

**15.6. Is  $(-\text{NH} - \text{CHR} - \text{CO}-)_n$  a homopolymer or copolymer?**

**Ans:** It is a homopolymer because the repeating structural unit has only one type of monomer, i.e.,  $\text{NH}_2 - \text{CHR} - \text{COOH}$ .

**15.7. In which classes, are the polymers classified on the basis of molecular forces?**

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**Ans:** Polymers are classified into four classes on the basis of molecular forces. These are:

elastomers, fibres, thermoplastic polymers and thermosetting polymers.

**1. Elastomers:** In these polymers, the intermolecular forces are the weakest. As a result, they can be readily stretched by applying small stress and regain their original shape when the stress is removed. The elasticity can be further increased by introducing some cross – links in the polymer chains. Natural rubber is the most popular example of elastomers. A few more examples are of: buna-S, buna-N and neoprene.

**2. Fibres:** Fibres represent a class of polymers which are thread-like and can be woven into fabrics in a number of ways. These are widely used for making clothes, nets, ropes, gauzes etc. Fibres possess high tensile strength because the chains possess strong intermolecular forces such as hydrogen bonding. These forces are also responsible for close packing of the chains. As a result, the fibres are crystalline in nature and have a sharp melting points. A few common polymers belonging to this class are nylon – 66, terylene and polyacrylonitrile etc.

**3. Thermoplastics:** These are linear polymers and have weak van der Waals forces acting in the various chains and are intermediate of the forces present in the elastomers and in the fibres. When heated, they melt and form a fluid which sets into a hard mass on cooling, Thus, they can be cast into different shapes by using suitable moulds. A few common examples are polyethene and polystyrene polyvinyls etc. These can be used for making toys, buckets, telephone apparatus, television cabinets etc.

**4. Thermosetting plastics:** These are normally semifluid substances with low molecular masses. When heated, they become hard and infusible due to the cross-linking between the polymer chains. As a result, they also become three dimensional in nature. They do not melt when heated. A few

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common thermosetting polymers are bakelite, melamine-formaldehyde, urea-formaldehyde and polyurethane etc.

**15.8. How can you differentiate between addition and condensation polymerisation?**

**Ans:** In addition polymerization, the molecules of the same or different monomers simply add on to one another leading to the formation of a macromolecules without elimination of small molecules like  $H_2O$ ,  $NH_3$  etc. Addition polymerization generally occurs among molecules containing double and triple bonds. For example, formation of polythene from ethene and neoprene from chloroprene, etc. In condensation polymerisation, two or more bifunctional trifunctional molecules undergo a series of independent condensation reactions usually with the elimination of simple molecules like water, alcohol, ammonia, carbon dioxide and hydrogen chloride to form a macromolecule. For example, nylon-6,6 is a condensation polymer of hexamethylenediamine and adipic acid formed by elimination of water molecules.

**15.9. Explain the term copolymerisation and give two examples.**

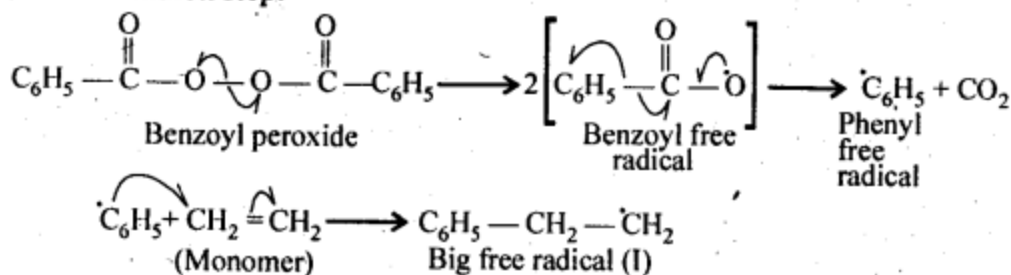
**Ans:** When two or more different monomers are allowed to polymerise together the product formed is called a copolymer, and the process is called copolymerisation. Example, Buna-S and Buna-N. Buna-S is a copolymer of 1, 3-butadiene and styrene while Buna-N is a copolymer of 1,3-butadiene and acrylonitrile.

**15.10. Write the free radical mechanism for the polymerisation of ethene.**

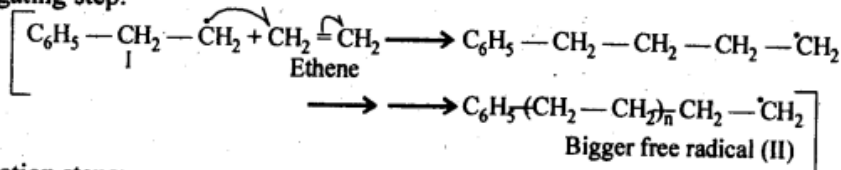
**Ans:**

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(a) Chain initiation step:

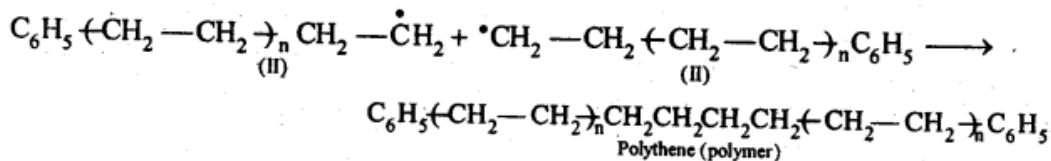


(b) Chain propagating step:

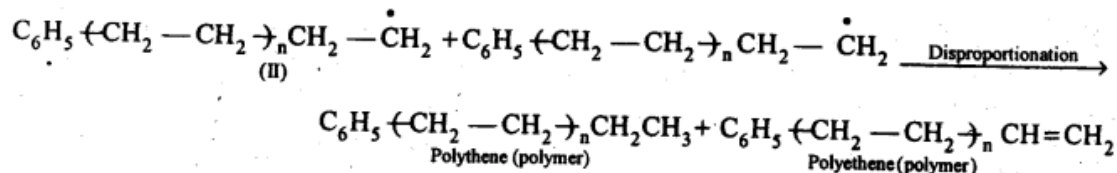


(c) Chain termination steps:

(i) By combination of free radicals (II):



(ii) By disproportionation of free radicals :



### 15.11. Define thermoplastics and thermo setting polymers with two examples of each

**Ans:** Thermoplastics polymers are linear polymer which can be repeatedly melted and moulded again and again on heating without any change in chemical composition and mechanical strength. Examples are polythene and polypropylene.

Thermosetting polymers, on the other hand, are permanently setting polymers. Once on heating in a mould, they get hardened and set, and then cannot be softened again. This hardening on heating is due to cross-linking

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between different polymeric chains to give a three dimensional network solid. Examples are bakelite, melamine-formaldehyde polymer etc.

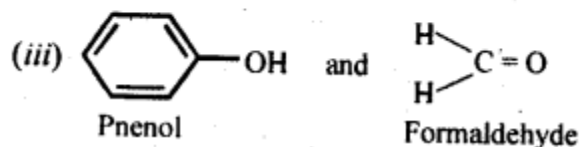
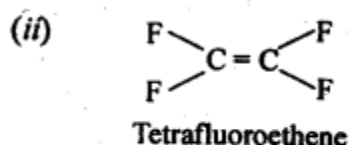
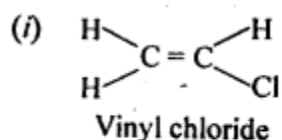
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**15.12. Write the monomers used for getting the following polymers:**

**(i) Polyvinylchloride**

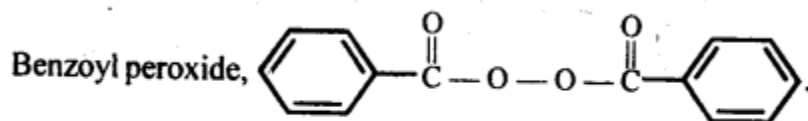
**(ii) Teflon (iii) Bakelite**

**Ans:**



**15.13. Write the name and structure of one of the common initiators used in free radical addition polymerisation.**

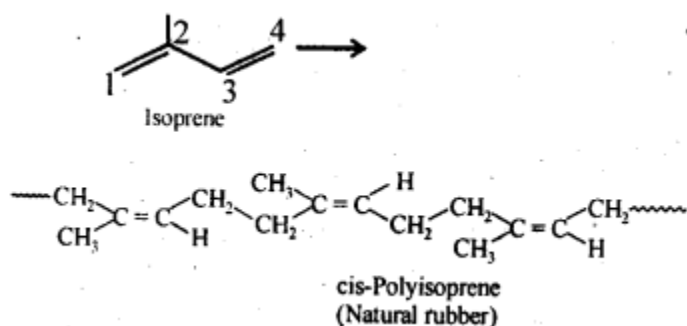
**Ans:**



**15.14. How does the presence of double bonds in rubber molecules influence their structure and reactivity?**

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**Ans:** Natural rubber is cis-polyisoprene and is obtained by 1,4-polymerization of isoprene units. In this polymer, double bonds are located between C<sub>2</sub> and C<sub>3</sub> of each isoprene unit. These cis-double bonds do not allow the polymer chains to come closer for effective interactions and hence intermolecular forces are quite weak. As a result, natural rubber, i.e., cis-polyisoprene has a randomly coiled structure not the linear one and hence show elasticity.



### 15.15. Discuss the main purpose of vulcanisation of rubber.

**Ans:** Natural rubber has the following disadvantages:

- (a) It is soft and sticky and becomes even more so at high temperatures and brittle at low temperatures. Therefore, rubber is generally used in a narrow temperature range (283-335 K) where its elasticity is maintained.
- (b) It has large water absorption capacity, has low tensile strength and low resistance to abrasion.
- (c) It is not resistant to the action of organic solvents.
- (d) It is easily attacked by oxygen and other oxidising agents. .

To improve all these properties, natural rubber is vulcanised by heating it with about 5% sulphur at 373-415 K. The vulcanized rubber thus obtained has excellent elasticity over a larger range of temperature, has low water

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absorption tendency and is resistant to the action of organic solvents and oxidising agents.

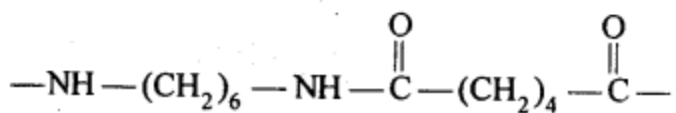
**15.16. What are the monomeric repeating units of Nylon-6 and Nylon 6,6?**

**Ans:**

The monomeric repeating unit of nylon-6

is  $\text{—NH—(CH}_2\text{)}_5\text{—}\overset{\text{O}}{\parallel}\text{C—}$  which is derived from caprolactam.

The monomeric repeating unit of nylon 6, 6 is derived from two monomers, hexamethylenediamine and adipic acid and has the following structure:

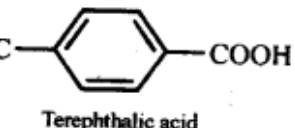


**15.17. Write the names and structures of the monomers of the following polymers:**

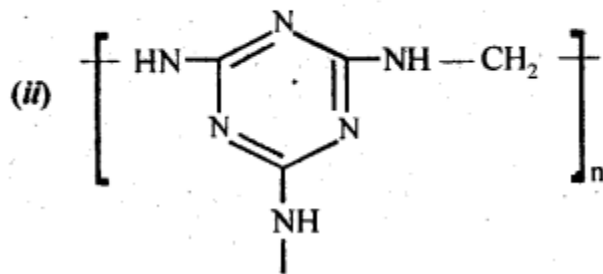
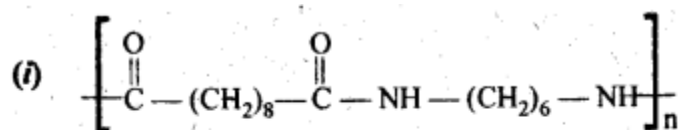
**(i) Buna-S (ii) Buna-N (iii) Dacron (iv) Neoprene**

**Ans:**

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- (i) Buna-S:  $\text{CH}_2 = \underset{\text{1,3-Butadiene}}{\text{CH} - \text{CH} = \text{CH}_2}$  and  $\text{C}_6\text{H}_5 - \underset{\text{Styrene}}{\text{CH} = \text{CH}_2}$ .
- (ii) Buna-N:  $\text{CH}_2 = \underset{\text{1,3-Butadiene}}{\text{CH} - \text{CH} = \text{CH}_2}$  and  $\text{CH}_2 = \underset{\text{Acrylonitrile}}{\text{CH} - \text{CN}}$ .
- (iii) Dacron:  $\text{HO} - \underset{\text{Ethylene glycol}}{\text{CH}_2 - \text{CH}_2} - \text{OH}$  and  $\text{HOOC} - \text{C}_6\text{H}_4 - \text{COOH}$ .  
  
 Terephthalic acid
- (iv) Neoprene:  $\text{CH}_2 = \overset{\text{Cl}}{\underset{|}{\text{C}}} - \text{CH} = \text{CH}_2$  and Chloroprene or 2-chloro-1,3-butadiene.

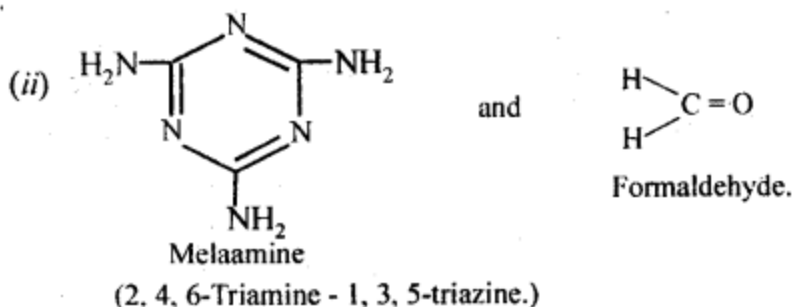
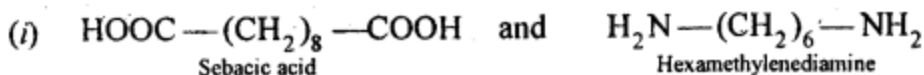
**15.18. Identify the monomer in the following polymeric structures:**



**Ans:**

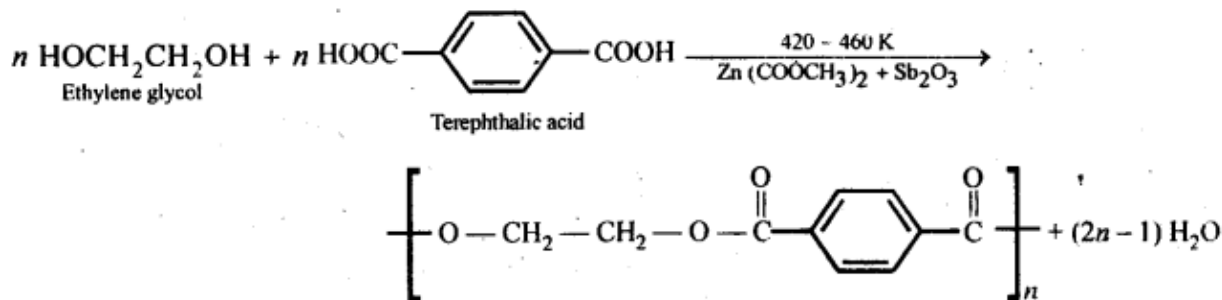
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**Monomers**



**15.19. How is dacron obtained from ethylene glycol and terephthalic acid?**

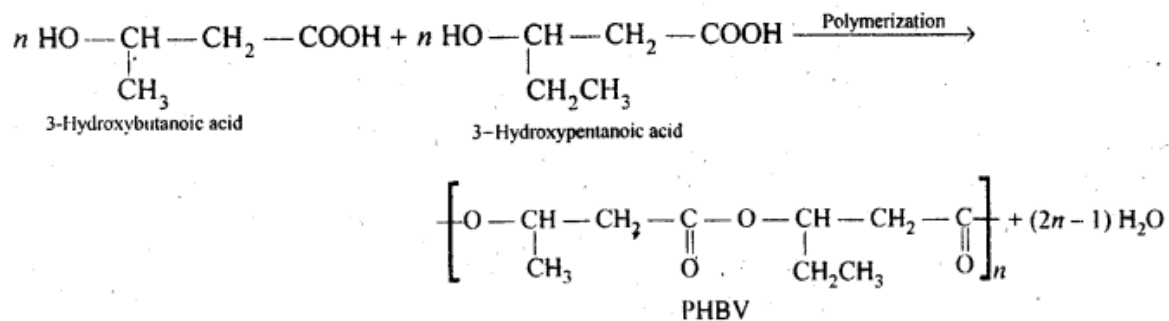
**Ans:** Dacron is obtained by condensation polymerization of ethylene glycol and terephthalic acid with the elimination of water molecules. The reaction is carried out at 420 – 460 K in presence of a catalyst consisting of a mixture of zinc acetate and antimony trioxide.



**15.20. What is a biodegradable polymer ? Give an example of a biodegradable aliphatic polyester.**

**Ans:** Polymers which disintegrate by themselves over a period of time due to environment degradation by bacteria, etc., are called biodegradable polymers. Example is PHBV, i. e., Poly-β-Hydroxybutyrate-co-β-Hydroxyvalerate.

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- Chapter 8 : The d and f  
Block Elements
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- Chapter 10 : Haloalkanes  
and Haloarenes
- Chapter 11 : Alcohols  
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Everyday Life

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