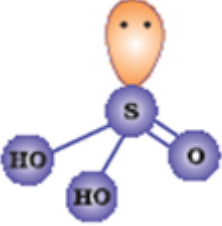
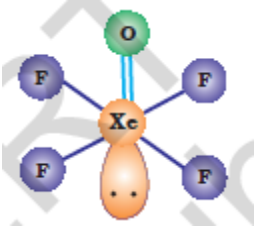
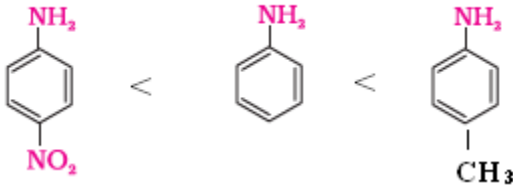
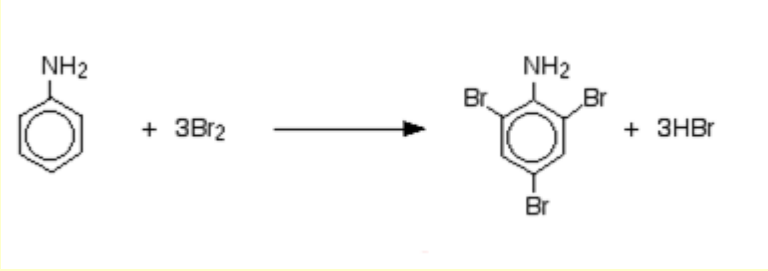
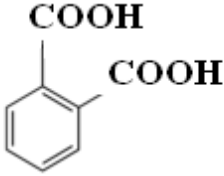
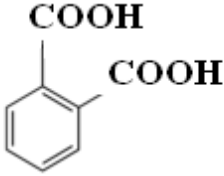
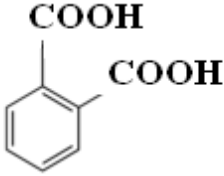
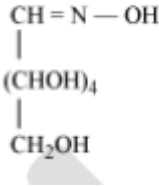
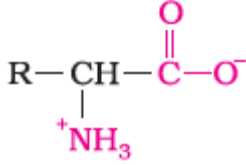
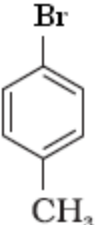


CHEMISTRY MARKING SCHEME
PATNA
SET -56/2/P

Qu es.	Answers	Marks
1	Because of no unpaired electron in Zn^{2+} Copper salts are coloured due to the presence of unpaired electrons in Cu^{2+}	$\frac{1}{2} + \frac{1}{2}$
2	$(\text{CH}_3)_3\text{C-Br}$	1
3	2F or 2x 96500C	1
4	Dispersed phase -liquid Dispersion medium - solid	$\frac{1}{2} + \frac{1}{2}$
5	2-Methylprop-2-en-1-ol	1
6	<div style="display: flex; justify-content: space-around; align-items: center;"> <div style="text-align: center;">  <p>(i)</p> </div> <div style="text-align: center;">  <p>(ii)</p> </div> </div>	1,1
7	Dichloridobis-(ethane-1,2-diamine)platinum(IV) Geometrical or optical isomerism OR (i) $[\text{Co}(\text{NH}_3)_6]\text{Cl}_3$ (ii) $\text{K}_2[\text{NiCl}_4]$	1 1 1 1
8	Decrease in concentration of reactant or increase in concentration of product per unit time Factors: 1)concentration of reactant 2)catalyst 3) temperature 4)Nature of reactant 5)pressure 6)surface area (any two)	1 $\frac{1}{2} + \frac{1}{2}$

9	<p>(i) $C_6H_5NH_2 < C_6H_5NHCH_3 < C_6H_5CH_2NH_2$</p> <p>(ii)</p> 	1 1
10	<p>Because on addition of a non volatile solute, vapour pressure of solution lowers down and therefore in order to boil solution, temperature has to be increased, thus boiling point gets higher</p> <p>Because it depends on molality/ number of solute particles / $\Delta T_b \propto m$</p>	1 1
11	<p>(i) Greater solubility of impurities in molten state.</p> <p>(ii) Silica reacts with impurity FeO to form slag ($FeSiO_3$) / acts as a flux to remove impurities.</p> <p>(iii) Cast iron is harder than pig iron / has lesser content of carbon.</p>	1 1 1
12	<p>(i) Because of the presence of triple bond between two N atoms / high bond dissociation enthalpy</p> <p>(ii) Because of the lowest bond dissociation enthalpy / least thermal stability.</p> <p>(iii) Because of low solubility in blood.</p>	1 1 1
13	<p>(i) $[CoF_6]^{3-}$ sp^3d^2 octahedral</p> <p>(ii) $[Ni(CN)_4]^{2-}$ dsp^2 square planar</p> <p>(b) CO , because of synergic /back bonding with metal</p>	$\frac{1}{2}$ $\frac{1}{2}$ $\frac{1}{2}$ $\frac{1}{2}$ $\frac{1}{2}$ $\frac{1}{2}$
14	<p>(i) $C_6H_5CONH_2 \xrightarrow{Br_2 + KOH} C_6H_5NH_2$</p> <p>(ii) $C_6H_5NH_2 \xrightarrow[0-5^\circ C]{NaNO_2 + HCl} C_6H_5N_2^+Cl^- \xrightarrow{H_2O} C_6H_5OH$</p> <p>(iii) $CH_3CN \xrightarrow{LiAlH_4} CH_3CH_2NH_2$</p> <p style="text-align: center;">OR</p>	1 1 1

14	<p>(i)</p>  <p>(ii)</p> $R-NH_2 + CHCl_3 + 3KOH \xrightarrow{\text{Heat}} R-NC + 3KCl + 3H_2O \quad (R = -C_6H_5)$ <p>(iii)</p> $C_6H_5NH_2 + HCl \longrightarrow C_6H_5NH_3^+ Cl^-$	1 1 1									
15	<table border="0"> <tr> <td data-bbox="129 846 448 943">i) Buna-S</td> <td data-bbox="448 846 879 943">Butadiene $CH_2=CH-CH=CH_2$</td> <td data-bbox="879 846 1425 943">Styrene $C_6H_5CH=CH_2$</td> </tr> <tr> <td data-bbox="129 943 448 1263">ii) Glyptal</td> <td data-bbox="448 943 879 1263">Ethylene Glycol $HO-CH_2CH_2-OH$</td> <td data-bbox="879 943 1425 1263">Phthalic acid </td> </tr> <tr> <td data-bbox="129 1263 448 1375">iii) Polyvinyl chloride</td> <td data-bbox="448 1263 879 1375">Vinyl Chloride</td> <td data-bbox="879 1263 1425 1375">$CH_2=CH-Cl$</td> </tr> </table> <p>(Note: half mark for name/s and half mark for structure/s)</p>	i) Buna-S	Butadiene $CH_2=CH-CH=CH_2$	Styrene $C_6H_5CH=CH_2$	ii) Glyptal	Ethylene Glycol $HO-CH_2CH_2-OH$	Phthalic acid 	iii) Polyvinyl chloride	Vinyl Chloride	$CH_2=CH-Cl$	$\frac{1}{2}$ $\frac{1}{2}$ $\frac{1}{2}$ $\frac{1}{2}$ $\frac{1}{2}$ $\frac{1}{2}$
i) Buna-S	Butadiene $CH_2=CH-CH=CH_2$	Styrene $C_6H_5CH=CH_2$									
ii) Glyptal	Ethylene Glycol $HO-CH_2CH_2-OH$	Phthalic acid 									
iii) Polyvinyl chloride	Vinyl Chloride	$CH_2=CH-Cl$									
16	<p>i)</p>  <p>(ii) Because of zwitter ion nature of amino acid /</p> <p>(iii) Because vitamin C is soluble in water.</p> 	1 1 1									

17	$\Delta T_f = K_f m$ $T_f^0 - T_f = \frac{K_f W_B \times 1000}{M_B \times W_A}$ $273K - T_f = 1.86K \text{ kg mol}^{-1} \times \frac{31g}{62g \text{ mol}^{-1}} \times \frac{1000}{500kg}$ $T_f = (273 - 1.86) K$ $T_f = 271.14K \quad \text{Or} \quad -1.86^\circ C$	1 1 1
18	(i) Unit cells having constituent particles at the corner positions. (ii) The defect occurs due to missing of equal no of cations and anions in a lattice. (iii) The permanent magnetism which arises when magnetic moments of substance are aligned in same direction.	1 1 1
19	i) $\text{CH}_3 - \text{CH}_2 - \overset{\text{Br}}{\underset{\text{CH}_3}{\text{C}}} - \text{CH}_3$ ii) $\text{CH}_3 - \text{CH}_2 - \text{CH} = \text{CH} - \text{CH}_3$ iii) 	1 1 1
20	(i) Because phenoxide ion is more stable than $\text{CH}_3\text{CH}_2\text{O}^-$ ion / due to resonance in phenol, oxygen acquires positive charge and releases H^+ ion easily whereas there is no resonance in $\text{CH}_3\text{CH}_2\text{OH}$ (ii) Because of hydrogen bonding in ethanol (iii) Because it follows SN_1 path way which results in the formation of stable $(\text{CH}_3)_3\text{C}^+$.	1 1 1
21	$\log \frac{K_2}{K_1} = \frac{E_a}{2.303R} \left[\frac{1}{T_1} - \frac{1}{T_2} \right]$ $\log \frac{4 \times 10^{-2}}{2 \times 10^{-2}} = \frac{E_a}{2.303 \times 8.314 \text{ J/K/mol}} \left[\frac{1}{300} - \frac{1}{310} \right]$ $\log 2 = \frac{E_a}{19.147 \text{ J/mol}} \left[\frac{10}{300 \times 310} \right]$ $E_a = \frac{0.3010 \times 19.147 \times 300 \times 310}{10}$ $E_a = 53598 \text{ J/mol} \quad \text{or} \quad 53.598 \text{ kJ/mol}$	1 1 1

22	(i) The zig-zag motion of the colloidal particles due to unbalanced bombardment by the particles of dispersion medium. (ii) The conversion of precipitate into colloidal sol by adding small amount of an electrolyte. (iii) On dissolution a large number of atoms or smaller molecules of a substance aggregate together to form species having size in the colloidal range.	1 1 1
23	i) Caring ,concerned, helping,empathy (any two) ii) By organizing competitions like slogan writing, poster making and talk in the morning assembly (any other correct answer) iii) Used to treat depression,. Iproniazid/phenelzine (any other correct example) iv) Saccharin/ sucralose/aspartame/alitame (any other correct example)	½ ½ 1 ½ ½ 1
24	$\text{Mg} \mid \text{Mg}^{2+} (0.001) \parallel \text{Cu}^{2+} (0.0001\text{M}) \mid \text{Cu}$ $E_{\text{cell}}^0 = E_{\text{R}}^0 - E_{\text{L}}^0$ $= [0.34 - (-2.37)]\text{V}$ $= 2.71\text{V}$ $E_{\text{cell}} = E_{\text{cell}}^0 - \frac{0.059}{n} \text{V} \log \frac{[\text{Mg}^{2+}]}{[\text{Cu}^{2+}]}$ $= 2.71\text{V} - \frac{0.059}{2} \text{V} \log 10^{-3}/10^{-4}$ $= 2.71 - 0.0295 \text{V} \log 10$ $= 2.71 - 0.0295$ $= 2.6805 \text{V}$ $\Delta G = -nFE_{\text{cell}}$ $= -2 \times 96500 \text{ C mol}^{-1} \times 2.68 \text{V}$ $= -517240 \text{ J mol}^{-1}$ $= -517.240 \text{ kJ/mol}$ <p style="text-align: center;">OR</p> a) $M = 0.20\text{M}$ $K = 2.48 \times 10^{-2} \text{S/cm}$ $\Lambda_m = \frac{K}{M} \times 1000 \text{ Scm}^2/\text{mol}$ $\Lambda_m = \frac{2.48 \times 10^{-2}}{0.20} \times 1000 \text{ Scm}^2/\text{mol}$ $= 124 \text{ Scm}^2/\text{mol}$ $\alpha = \frac{\Lambda_m}{\Lambda_m^0}$ $\Lambda_m^0 = \lambda^0 \text{K}^+ + \lambda \text{Cl}^-$	1 1 1 1 ½ ½ 1 ½
24	a) $M = 0.20\text{M}$ $K = 2.48 \times 10^{-2} \text{S/cm}$ $\Lambda_m = \frac{K}{M} \times 1000 \text{ Scm}^2/\text{mol}$ $\Lambda_m = \frac{2.48 \times 10^{-2}}{0.20} \times 1000 \text{ Scm}^2/\text{mol}$ $= 124 \text{ Scm}^2/\text{mol}$ $\alpha = \frac{\Lambda_m}{\Lambda_m^0}$ $\Lambda_m^0 = \lambda^0 \text{K}^+ + \lambda \text{Cl}^-$	½ 1 ½

	$= 73.5 + 76.5$ $= 150.5$ $\alpha = \frac{124}{150} = 0.82$ Or 82%	1
	b) Primary battery or cell, potential remains constant throughout its life.	1,1
25	a) <ul style="list-style-type: none"> i) Due to lanthanoid contraction. ii) Due to incomplete filling of d- orbitals / comparable energies of (n-1)d & ns electrons. iii) Because it undergoes disproportionation reaction in aqueous solution/ oxidation of a metal in a solvent depends on the nature of the solvent. Cu^+ is unstable in water that's why it undergoes oxidation. b) <ul style="list-style-type: none"> i) $2\text{MnO}_2 + 4\text{KOH} + \text{O}_2 \rightarrow 2\text{K}_2\text{MnO}_4 + 2\text{H}_2\text{O}$ ii) $2\text{Na}_2\text{CrO}_4 + 2\text{H}^+ \rightarrow \text{Na}_2\text{Cr}_2\text{O}_7 + \text{H}_2\text{O} + 2\text{Na}^+$ <p style="text-align: center;">OR</p>	1 1 1
25	a) <ul style="list-style-type: none"> (i) Because of high $\Delta_a\text{H}^\circ$ & low $\Delta_{\text{hyd}} \text{H}^\circ$. (ii) Because of more stability of Mn^{2+} ($3d^5$) (iii) Cr^{2+}, because in +3 oxidation state Cr is more stable (t^3_{2g} orbital) b) Due to comparable energies of 5f, 6d, 7s orbitals. Both show contraction in size/ both show main oxidation state +3/both are electro positive and very reactive/ both exhibit magnetic and spectral properties. (any one)	1 1 $\frac{1}{2}, \frac{1}{2}$ 1 1

26	<p>a) CH_3COCl (A) CH_3CHO (B) $\begin{array}{c} \text{OH} \\ \\ \text{CH}_3\text{CH}-\text{CH}_2-\text{CHO} \end{array}$ (C) $\text{CH}_3\text{CH}=\text{CH}-\text{CHO}$ (D)</p> <p>b) i) On adding Tollen's reagent $\text{C}_6\text{H}_5\text{CHO}$ forms silver mirror whereas $\text{C}_6\text{H}_5\text{COCH}_3$ does not.</p> <p>ii) On adding NaHCO_3 solution benzoic acid gives brisk effervescence but methyl benzoate does not.</p> <p>(or any other distinguishing test)</p> <p>c) $\begin{array}{c} \text{CH}_3\text{CH}_2-\text{CH}-\text{CHO} \\ \\ \text{CH}_3 \end{array}$</p>	<p>$\frac{1}{2}$, $\frac{1}{2}$</p> <p>$\frac{1}{2}$, $\frac{1}{2}$</p> <p>1</p> <p>1</p> <p>1</p>
26	<p style="text-align: center;">OR</p> <p>a) i) $\text{CH}_3\text{CH}_2\text{CH}_3$</p> <p>ii) $\begin{array}{c} \text{CH}_3-\text{C}=\text{N}-\text{NHCONH}_2 \\ \\ \text{CH}_3 \end{array}$</p> <p>iii) $\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3-\text{C}-\text{OH} \\ \\ \text{CH}_3 \end{array}$</p> <p>b) $\text{CH}_3\text{CHO} < \text{CH}_3\text{CH}_2\text{OH} < \text{CH}_3\text{COOH}$</p> <p>c) On adding Tollen's reagent $\text{CH}_3\text{CH}_2\text{CHO}$ forms silver mirror whereas $\text{CH}_3\text{CH}_2\text{COCH}_3$ does not (or any other distinguishing test).</p>	<p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>1</p>