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KN-373

D. Pharmacy (Part-I) Examination-2022

PHARMACEUTICAL CHEMISTRY-I

[Paepr : Second]

Time Allowed: Three Hours

Maximum Marks: 80

Minimum Marks: 27

(Section-A)

Objective Type Questions

Q.I Attempt all questions. 20 x1 = 20

1. What is calamine?

a) Zinc oxide

b) Zinc oxide with traces of magnesium oxide

c) Zinc oxide with traces of ferric oxide

d) Zinc carbonate

2. Which one of these is a major cation in intracellular fluid?

a) Potassium

b) Sodium

c) magnesium

d) calcium

3. Titrant used in the assay of ferrous sulphate I.P.1996 is-

a) Potassium parmangnate

b) cerric ammonium sulphate

c) sulphuric acid

d) None of the above

4. The lead acetate cotton wool is used in limit test for arsenic to-

a) Make the arsine gas

b) Develop the yellow colour

c) Trap the hydrogen sulphide gas

d) None of the above

5. Limit tests are performed in-

a) Flask

b) Test Tube

c) Nesslar cylinder

d) None of the above

6. Silver nitrate should be stored in-

a) Clean, dry white bottles

b) Lead free white bottle

c) Amber colourd bottle

d) Neutral glass bottle

7. The drug commonly used to prevent sepsis in severe burn cases is-

a) Sulfamethoxazole

b) Silversulphadiazine

c) Salazopyrine

d) Sulfadiamidine

8. Which of the following is used as cryotherepeutic agent?

a) Calcium carbide

b) Liquid nitrogen

c) Calcium cyanamide

d) Sodium nitrite

9. which of the following is used in the eye treatment in minor infections?

a) Ammoniated mercury

b) Yellow mercuric oxide

c) Potassium permangnate

d) None of the above

10. Which of the following is used as Rust inhibitor?

a) Sodium nitrite

b) Liquid nitrogen

c) Calcium carbide

d) Sodium thiosulphate

11. Antidote used in Cyanide poisoning is-

a) Sodium chloride

b) Potassium iodide

c) Sodium thiosulphate

d) Sodium bicarbonate

12. Unit of measurement of radioactivity is-

a) Curie

b) Cm

c) Litre

d) Gm

13. Which of the following particles have no Charge and mass?

a) α- particles

b) β-particles

c) γ-particles

d) All of the above

14. Milk of magnesia is a common name for-

a) Suspension of magnesium oxide

b) suspension of magnesium hydroxide

c) suspension of magnesium carbonate

d) None of the above

15. Epsom salt is-

a) Sodium sulphate

b) Potassium sulphate

c) Magnesium sulphate

d) None of the above

16. Hydrochloric Acid (HCl) is used in-

a) Diarrhea

b) Hyperacidity

c) Achlorhydria

d) Fever

17. The colour of the cylinder of nitrous oxide is-

a) Red

b) Blue

c) Black

d) Orange

18. A gas used for fumigation is-

a) Nitrogen oxide

b) Ethylene

c) Oxygen

d) Sulphur dioxide

19. Antimony potassium tartarate is used as-

a) Antidote

b) Emetic

c) Astringent

d) Buffer

20. Iodine dissolves in water with the addition of-

a) Sodium chloride
b) Sulphur
c) Potassium iodide
d) Calcium carbonate

(Section-B)

Short Answer Type Questions

Q.II Attempt any ten question. 10 x3 = 30

1. Discuss importance of Radioisotopes in pharmacy.
2. Define different terminologies as antimicrobial agents.
3. Explain Lewis acid-base theory with examples.
4. Define and classify laxatives with examples.
5. Explain G.I.T. protective adsorbent. Give properties and uses of kaolin.
6. Discuss Bronsted-Lowry concept of acids and bases. Explain its advantages over Arrhenius theory.

7. Explain ORS powders recommended by UNICEF and WHO.
8. What are antacids? Give important properties of antacid.
9. Define antidote. Discuss various actions of antidotes with examples.
10. Explain the role of iron in human body.
11. Give any four properties of α and β rays.

(Section-C)

Long Answer Type Questions

Q.III Attempt any six questions 6 x5 = 30

1. Explain Pharmaceutical Buffer Solution with importance.
2. Discuss about physiological acid – base balance.
3. Write limit test for chloride.

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4. Write notes on Protective & Adsorbents.
5. Give Exhaustive notes on Astringents.
6. What are Metabolic acidosis and alkalosis? How they are treated?
7. Define respiratory stimulant and expectorant. State properties of potassium iodide.
